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Holt ChemFile: Problem-Solving Workbook 55 Mole Concept. Name Class Date. Problem Solving continued. CONVERTING THE AMOUNT OF AN ELEMENT IN MOLES TO THE NUMBER OF ATOMS In Sample Problem 3, you were asked to determine the number of moles in 1.204 10<sup>24</sup> atoms of lithium.

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## **Holt Chemfile Problem Solving Workbook Mole Concept Answers**

Holt ChemFile: Problem-Solving Workbook 58 Mole Concept Name Class Date Problem Solving continued CONVERTING NUMBER OF ATOMS OF AN ELEMENT TO MASS Sample Problem 4 uses the progression of steps 1→2→3 to convert from the mass of an element to the number of atoms. In order to calculate the mass from a given number of atoms, these steps will be reversed.

## **Skills Worksheet Problem Solving**

Holt ChemFile: Problem-Solving Workbook 5 Conversions Name Class Date Problem Solving continued Sample Problem 2 A metallurgist is going to make an experimental alloy that requires adding 325 g of bismuth to 2.500 kg of molten lead. What is the total mass of the mixture in kilograms? Solution ANALYZE

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Holt ChemFile: Problem-Solving Workbook 259 pH. pH. In 1909, Danish biochemist S. P. L Sørensen introduced a system in which acidity was expressed as the negative logarithm of the H concentration. In this way, the acidity of a solution having H concentration of  $10^{-5}$  M would have a value of 5.

## Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook 99 Stoichiometry  
Name Class Date Problem Solving continued Sample Problem 1  
Ammonia is made industrially by reacting nitrogen and hydrogen under pressure, at high temperature, and in the presence of a catalyst. The equation is  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$ . If 4.0 mol of  $H_2$  react, how many moles of  $NH_3$  will be produced?

## Skills Worksheet Problem Solving

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## Holt Chemfile Problem Solving Workbook Answers Stoichiometry

Holt ChemFile: Problem-Solving Workbook 198 Concentration of Solutions Problem Solving continued Sample Problem 1 What is the percentage by mass of a solution made by dissolving 0.49 g of potassium sulfate in 12.70 g of water? Solution ANALYZE What is given in the problem? What are you asked to find? Items Data  
Mass of solvent 12.70 g  $H_2O$

## Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook 1 Conversions  
Conversions One of the aims of chemistry is to describe changes—to tell what changed, how it changed, and what it changed into. Another aim of chemistry is to look at matter and its changes and to ask questions such as how much, how big, how hot, how many, how hard, and how long did it take.

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## **Solving ...**

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Ammonia is made industrially by reacting nitrogen and hydrogen under pressure, at high temperature, and in the presence of a catalyst. The equation is  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$ .

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Name Class Date Problem Solving continued CONVERTING  
NUMBER OF ATOMS OF AN ELEMENT TO MASS Sample Problem 4  
uses the progression of steps 1 → 2 → 3 to convert from the mass of an element to the number of atoms.

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Name Class Date • Problem Solving continued PROBLEMS  
INVOLVING ATOMS AND ELEMENTS Sample Problem 1 A chemist  
has a jar containing 388.2 g of iron filings. How many moles of  
iron does the jar contain? Solution ANALYZE What is given in the  
problem? mass of iron in grams

**continued - PC\|MAC**

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